Singlet and Triplet States Are Degenerate in 2,3-Dimethylenecyclohexane-1,4-diyl

Kenji Matsuda and Hiizu Iwamura*

Institute for Fundamental Research in Organic Chemistry Kyushu University, 6-10-1 Hakozaki Higashi-ku, Fukuoka, 812-81 Japan

Received March 17, 1997 Revised Manuscript Received May 27, 1997

Trimethylenemethane (TMM) and tetramethyleneethane (TME) are the simplest non-Kekulé alternant hydrocarbons (AHs) for which the Longuet-Higgins rule dictates the occurrence of two nonbonding molecular orbitals (NBMO) that have zero π -bond energy.¹ The two NBMOs have atoms in common in TMM and can be confined to different sets of atoms in TME, and Hund's rule can not therefore be applicable in a straightforward manner to the latter. Pointing out this difference, Borden and Davidson predicted in their perturbational MO theory that while TMM should have a ground triplet state, singlet and triplet states are nearly degenerate and higher order terms favor the singlet ground state for TME.² Ovchinnikov adopted valence bond (VB) theory for AHs to conclude that the spin quantum number (S) of the AHs can be given by $S = (n^* - n)/2$, where n^* and n are the numbers of starred and unstarred carbon atoms in AHs, respectively. The S values of TMM and TME are predicted by this rule to be 1 and 0, respectively.³

While the triplet ground state of TMM was accepted unanimously,⁴ it was only in 1986 when the ground spin state of TME was tested experimentally by Dowd *et al.* They studied the ESR spectra of TME and 2,3-dimethylenecyclohexane-1,4diyl (**1**), a conformationally restricted analog, to find that their signal intensities obeyed Curie law in the temperature ranges of $16-65^{5d}$ and 15-53 K,^{5f} respectively. While the possibility that the singlet and triplet states are degenerate within 30-40cal mol⁻¹ was mentioned, their triplet ground states were concluded on the basis that it was clearly less likely that both TME and **1** of different planarity showed similar degeneracy. This conclusion was contradictory to the above theoretical predictions and the observed singlet ground states of higher vinylogs and benzenologs of TME.⁶ Recently Berson *et al.* investigated *N*-substituted-3,4-dimethylenepyrroles extensively and reported that while *N*-methyl- or *N*-pivaloyl-3,4-dimethylenepyrroles were characterized as singlet biradicals, *N*-tosylor *N*-brosyl-3,4-dimethylenepyrroles existed as both singlet and triplet biradicals and intersystem crossing between the two spin states could not be observed. They concluded substitution by an electron-withdrawing group at the nitrogen of the pyrrole ring can tune the singlet—triplet energy gap.⁷

A question how the singlet and triplet ground states of TME are separated evoked much controversy and should be answered by proper methods.⁸ Determination of temperature dependence of the magnetic susceptibility and magnetic field dependence of the magnetization is a method of choice for studying the magnitude of the effective magnetic moments and spin quantum numbers of polyradicals.^{6,9} In this paper, we report a magnetic study directed to determine the energy gap of the singlet and triplet states of **1**.

Diradical **1** was obtained by the photolysis of 5,6-dimethylene-2,3-diazabicyclo[2.2.2]oct-2-ene (**2**) (eq 1) as employed by Dowd *et al.*⁴ Precursor **2** was synthesized from 1,2-dihydrophthalic acid according to the method of Roth *et al.*¹⁰ and obtained as pale yellow crystals {UV-vis absorption maxima at 274 nm (ϵ 4990) and 379 nm (ϵ 112)}. Since the magnetic



measurement is not spectroscopic and treats the photoproducts as a whole, it is necessary to minimize any side reaction giving paramagnetic impurities.⁹ We therefore optimized the reaction conditions for the formation of **1** from **2** in 2-methyltetrahydrofuran (MTHF) frozen solution. The photolysis was carried out in an ESR cavity by using light from a high-pressure mercury lamp and a set of Kenko band-path filters. While the light $\lambda > 340$ nm was not effective in giving any ESR active product, the light ($\lambda > 320$ nm) gave g = 2 signals (|D/hc| = 0.024 cm⁻¹ and |E/hc| = 0.0039 cm⁻¹) as well as forbidden lines at g = 4 (Figure 1a). The *D* and *E* values correspond to the theoretical values^{11,12} for **1** and are in good agreement with a distance of 4.77 Å between the two spins by a point-dipole approximation. We conclude this species as the desired diradical **1**.

(11) (a) Rule, M.; Matlin, A. R.; Seeger, D. E.; Hilinski, E. F.; Dougherty,
D. A.; Berson, J. A. *Tetrahedron* 1982, 38, 787. (b) Prasad, B. L. V.;
Radhakrishnan, T. P. J. Mol. Struct. (THEOCHEM) 1996, 361, 175.

(12) A CASSCF(6,6)/6-31G* calculation for 1 at the B3LYP(6-311++G**) geometry (23° twist) has been performed by Havlas and Michl to yield $|D/hc| = 0.0329 \text{ cm}^{-1}$ and $|E/hc| = 0.0180 \text{ cm}^{-1}$. We thank Prof. Josef Michl for the information prior to the publication.

 ^{(1) (}a) Longuet-Higgins, H. C. J. Chem. Phys. **1950**, 18, 265. (b) Dowd,
 P. Acc. Chem. Res. **1972**, 5, 242. (c) Salem, L.; Rowland, C. Angew. Chem., Int. Ed. Engl. **1972**, 11, 92. (d) Berson J. A. Acc. Chem. Res. **1978**, 11, 446. (e) Borden, W. T., Ed.; Diradicals; Wiley: New York, 1982. (f) Lahti,
 P. M.; Ichimura, A. S.; Berson, J. A. J. Org. Chem. **1989**, 54, 958. (g) Berson, J. A. Mol. Cryst. Liq. Cryst. **1989**, 176, 1. (h) Borden, W. T.; Davidson, E. R. Acc. Chem. Res. **1996**, 29, 67.

⁽²⁾ Borden, W. T.; Davidson, E. R. J. Am. Chem. Soc. 1977, 99, 4587.
(3) Ovchinnikov, A. A. Theor. Chim. Acta 1978, 47, 297.

^{(4) (}a) Dowd, P. J. Am. Chem. Soc. 1966, 88, 2587. (b) Crawford, R. J.;
Cameron D. M. J. Am. Chem. Soc. 1966, 88, 2589. (c) Dowd, P.; Sachdev, K. J. Am. Chem. Soc. 1967, 89, 715. (d) Dowd, P.; Gold, A.; Sachdev, K. J. Am. Chem. Soc. 1968, 90, 2715. (e) Berson, J. A.; Bushby, R. J.; Tremelling, M.; McBride, J. M. J. Am. Chem. Soc. 1971, 93, 1544. (f) Carpenter, B. K.; Little, R. D.; Berson, J. A. J. Am. Chem. Soc. 1971, 93, 5723. (g) Platz, M. S.; McBride, J. M.; Little, R. D.; Harrison, J. J.; Shaw, A.; Potter, S. E.; Berson, J. A. J. Am. Chem. Soc. 1976, 98, 5725. (h) Baseman, R. J.; Pratt, D. W.; Chow, M.; Dowd, P. J. Am. Chem. Soc. 1976, 98, 5726.

<sup>98, 5/26.
(5) (</sup>a) Dowd, P. J. Am. Chem. Soc. 1970, 92, 1066. (b) Roth, W. R.;
Erker, G. Angew. Chem., Int. Ed. Engl. 1973, 12, 503. (c) Beetz, T.; Kellogg,
R. M. J. Am. Chem. Soc. 1973, 95, 23. (d) Dowd, P.; Chang, W.; Paik, Y.
H. J. Am. Chem. Soc. 1986, 108, 7416. (e) Roth, W. R.; Kowalczik, U.;
Maier, G.; Reisenauer, H. P.; Sustmann, R.; Müller, W. Angew. Chem.,
Int. Ed. Engl. 1987, 26, 1285. (f) Dowd, P.; Chang, W.; Paik, Y. H. J. Am.
Chem. Soc. 1987, 109, 5284. (g) Roth, W. R.; Langer, R.; Bartmann, M.;
Stevermann, B.; Maier, G.; Reisenauer, H. P.; Sustmann, R.; Müller, W.
Angew. Chem., Int. Ed. Engl. 1987, 26, 256.

^{(6) (}a) Matsumoto, T.; Koga, N.; Iwamura, H. J. Am Chem. Soc. **1992**, 114, 5448. (b) Matsumoto, T.; Ishida, T.; Koga, N.; Iwamura, H. J. Am. Chem. Soc. **1992**, 114, 9952. (c) Borden, W. T.; Iwamura, H.; Berson, J. A. Acc. Chem. Res. **1994**, 27, 109.

^{(7) (}a) Bush, L. C.; Heath, R. B.; Berson, J. A. J. Am. Chem. Soc. 1993, 115, 9830. (b) Bush, L. C.; Heath, R. B.; Feng, X. W.; Wang, P. A.; Maksimovic, L.; Song, A. I.; Chung, W.-S.; Berinstain, A. B.; Scaiano, J. C.; Berson, J. A. J. Am. Chem. Soc. 1997, 119, 1406. (c) Bush, L. C.; Maksimovic, L.; Feng, X. W.; Lu, H. S. M.; Berson, J. A. J. Am. Chem. Soc. 1997, 119, 1416.

^{(8) (}a) Du, P.; Hrovat, D. A.; Borden, W. T. J. Am. Chem. Soc. 1986, 108, 8086. (b) Du, P.; Borden, W. T. J. Am. Chem. Soc. 1987, 109, 930.
(c) Radhakrishnan, T. P. Chem. Phys. Lett. 1991, 181, 455. (d) Prasad, B. L. V.; Radhakrishnan, T. P. J. Phys. Chem. 1992, 96, 9232. (e) Nachtigall, P.; Jordan, K. D. J. Am. Chem. Soc. 1992, 114, 4743. (f) Nash, J. J.; Dowd, P.; Jordan, K. D. 1992, 114, 10071. (g) Pranata, J. J. Am. Chem. Soc. 1992, 114, 10537. (h) Nachtigall, P.; Jordan, K. D. J. Am. Chem. Soc. 1993, 115, 270.

⁽⁹⁾ Iwamura, H.; Koga, N. Acc. Chem. Res. 1993, 26, 346.

⁽¹⁰⁾ Roth, W. R.; Biermann, M.; Erker, G.; Jelich, K. Chem. Ber. 1980, 113, 586.



Figure 1. (a) A 9.428 GHz ESR spectrum of a solid solution of **2** in MTHF (0.1 M) after irradiation with light ($\lambda > 320$ nm) measured at 9.4 K. The signal denoted with an asterisk (*) appeared only when irradiated with light ($\lambda > 300$ nm). The $\Delta m_s = \pm 2$ region is also presented. (b) An ESR spectrum obtained by continued irradiation at $\lambda > 300$ nm at 9.5 K. The signal denoted with Δ was due to the impurity in the cavity.

hotolysis of **2** or continued irradiation of **1** with the light λ > 300 nm gave another set of ESR signals (Figure 1b). The second photoproduct appeared at the expense of the signal from the first product. If the outer-to-inner peaks and shoulders of these signals are assumed to be the $\pm z$, $\pm y$, and $\pm x$ transitions of a triplet species, $|D/hc| = 0.0059 \text{ cm}^{-1}$ and $|E/hc| = 0.00067 \text{ cm}^{-1}$ are obtained. However, no forbidden transition is observed at g = 4, proven by the decrease of the signal at g = 4. The small *D* and *E* values, corresponding to the distance between the two spins of 7.61 Å by a point-dipole approximation, may be assinged to a hydrogen migration product. The ESR spectrum reported by Dowd *et al.*^{5f} is reproduced by a sum of the two species.

The wavelength-dependence of the photoproduct has been reported by Berson *et al.* in the *N*-(arenesulfonyl)-3,4-dimethylenepyrrole system.⁷ The irradiation at 370 nm afforded the ESR inactive singlet biradical, but irradiation at 265 nm afforded a triplet species. With irradiation at 343 nm, a different triplet species (smaller *D* and *E* values) was produced. Both triplet species exhibited a forbidden line at g = 4. The triplet species with larger *D* and *E* values was assigned to a triplet biradical, and the triplet species with the smaller *D* and *E* values was assigned to the triplet diazenyl biradical.

Magnetic measurements were performed on a SQUID susceptometer/magnetometer for samples of 1 obtained under similar conditions as those described above. A solution of 0.33 mg of 2 in 10 μ L of MTHF was placed in a quartz cell. The light ($\lambda > 320$ nm) was introduced to the SQUID susceptometer through an optical fiber introduced into a sample probe and the photolysis of 2 in MTHF was performed at 4–10 K. The data before irradiation was subtracted as background data.

The paramagnetic susceptibility χ obtained at a field of 5000 G in the temperature range 2–20 K is presented as a χT vs T plot (Figure 2). The nearly horizontal line is consistent with Dowd's finding that the EPR signal intensity of **1** obeys the Curie law. The slightly decrease in the χT values at the low-temperature region (T < 4 K) suggests a singlet-triplet equilibrium in which the singlet is very slightly lower in energy, rather than a triplet ground state. The effect is, however, too small to apply a model in which a ground singlet state is in equilibrium with a triplet separated by a small gap (2*J*). Nevertheless, this theoretical curve for $2J/k_B = -0.76$ K derived from the field dependence of magnetization (*vide infra*) is also given in Figure 2.



Figure 2. χT vs *T* plots of **1** measured at 5000 G. The ordinate is based on the amount of **2** placed in a sample cell and not normalized on the basis of the amount of **1** produced. The solid curve is a theoretical one for a singlet-triplet model with a gap of $2J/k_{\rm B} = -0.76$ K.



Figure 3. Field dependence of the magnetization of **1** measured at 2.0 K. The two solid curves represent a theoretical S = 1 curve (top) and a best-fit curve (bottom) with a singlet—triplet gap of $2J/k_{\rm B} = -0.76$ K (eq 2).

Field-dependence of the magnetization (M) of **1** is shown in Figure 3. A theoretical formula for the magnetization of a diradical with a small singlet and triplet energy gap (2J) comparable to the Zeeman splitting is given by eq 2. The fitting

$$M = Ng\mu_{\rm B} \frac{\exp\left(\frac{g\mu_{\rm B}H_0}{k_{\rm B}T}\right) - \exp\left(\frac{-g\mu_{\rm B}H_0}{k_{\rm B}T}\right)}{\exp\left(\frac{g\mu_{\rm B}H_0}{k_{\rm B}T}\right) + 1 + \exp\left(\frac{-g\mu_{\rm B}H_0}{k_{\rm B}T}\right) + \exp\left(\frac{-2J}{k_{\rm B}T}\right)}$$
(2)

of eq 2 to the observed data gave $2J/k_{\rm B} = -0.76 \pm 0.16$ K, indicating that the singlet and triplet states are nearly degenerate in **1**. For a triplet (S = 1) ground state ($J \gg 0$), the last term in the denominator vanishes and the corresponding theoretical curve did not fit the observed data at all. For $J \approx 0$, this term is 1 and, therefore, the initial slope of the curve for $J \approx 0$ is three-fourths that of the curve for $J \gg 0$, indicating that 25% of the diradical molecules are in the singlet state and hence give no magnetization.¹³

These experimental results are in excellent accord with some of the computational results that the singlet-triplet gap in 1 is almost exactly zero.⁸

In conclusion, we have been able to demonstrate experimentally that the singlet and triplet states of 1, a conformationally fixed analog of TME, are almost degenerate, terminating a controversy as to theoretical and experimental interpretation of the ground state of 1.

Acknowledgment. This work supported by a Grant-in-Aid for Scientific Research (B) (no. 06453035) from the Ministry of Education, Science, Sports, and Culture, Japan.

JA970843Z

⁽¹³⁾ The theoretical S = 1 curve is normalized at 50 kOe. The initial slope of the curve is therefore nearly but not quite as high as four-thirds that of the curve for $2J/k_{\rm B} = -0.76$ K.